Species Tag: Version: Date: Contributor:	40002 1 Aug. 1987 R. L. Poynter	Species Name:	NaOH Sodium hydroxide
Lines Listed: Freq. (GHz) < Max. J: LOGSTR0= LOGSTR1= Isotope Corr.: Egy. (cm <sup>-1</sup> ) > $\mu_a = \mu_b = \mu_c = \mu_c = \mu_c$	91 2201 91 -6.3 -8.0 0.0 0.0 1.0	Q(300.0) = Q(225.0) = Q(150.0) = Q(75.00) = Q(37.50) = Q(18.75) = Q(9.375) = A = B = C = C = Q(300.0) = Q(30	498.884 374.025 249.345 124.767 62.532 31.427 15.882 12567.0590(120)

The measured lines were taken from E. F. Pearson and M. B. Trueblood, 1973, Astrophys. J. Lett. Ed. 179, L146, and P. Kuipers et al., 1976, Chem. Phys. 15, 457. Because no dipole moment measurement has been reported, we assumed a value of unity in our calculation. Two theoretical estimates have been published: N. G. Rambidi, Yu. G. Abashkin, and A. I. Dement'ev, 1984, Russ. J. Inorg. Chem. 29, 12–22, and G. A. Long, J. F. Capitani, and L. Pedersen, 1983, J. Mol. Struct. 105, 229–230. Of these, the Rambidi et al. value is closer to our estimate of 6.63 Debye with an estimated error of maybe 1 Debye, which is based upon extrapolations from the dipole moments of Li, Cs, and Na fluorides and the Cs and Li hydroxides. The Rambidi et al. value is 6.706 debye, while the Long et al. value ranges between 6.30 and 5.42 debye, depending on the basis set used in their calculations. The line intensities should be multiplied by a factor of about 44 to get something closer to the right value. The theoretical calculations suggest that this molecule is linear. A measurement is needed.